Year 10 Chemistry WINTER

I CAN:	☺	☹
SAFETY		
Recognise hazards in the laboratory		
Recognise hazard symbols and their meanings		
Physical changes		
Name the changes of state		
Explain what happens in terms of particles during a change of state		
Describe the properties of solids, liquids and gases		
Draw diagrams of the arrangement of particles in solids, liquids and gases		
Chemical changes		
Describe evidence of a chemical change		
State that during a chemical change a new substance is made		
Recall the differences between a chemical and a physical change		
Atomic structure		
Recall the charge, mass and position of the particles found in an atom		
Define atomic number and mass number		
Draw an atom of an element (atomic numbers 1- 20)		
Define an Isotope		
Define an Ion and be able to draw the electronic structure of an ion		
Ionic Bonding		
Recall that Ionic bonding occurs between a metal and a non-metal		
Recall that Electrostatic force holds oppositely charged particles together		
Recall the term for a positive ion and a negative ion		
Draw the electronic structure of an ionic compound (eg. NaCl) including square brackets and charges		

Covalent bonding		
Recall definition of a single covalent bond		
Be able to draw a dot and cross diagram of molecules of H_2 H_2O, HCl		
Metallic bonding		
Draw a diagram to represent metallic bonding, showing a lattice of metal ions surrounded by a sea of mobile electrons.		